**Revision Sheet #3** **|** Lower Secondary Stage (6-8)

1st Semester | 2023-2024

 **Name:** \_\_\_\_\_\_\_\_\_\_\_

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| --- | --- |
| **Subject:** Science/ Chemistry | **Chapter:** 4.8 Understanding Salts 4.9 Exploring reactions of acids with metals |
| **Objectives:*** Describe the use of common salts.
* Predict salts made in neurtalisation reactions.
* Describe reactions between acids and metals.
 |

**1 – Complete the table below with information about the salts listed.**

|  |  |  |  |
| --- | --- | --- | --- |
| **Salt** | **Chemical formula** | **Which acid and alkali can produce it** | **Uses** |
| Sodium chloride |  |  |  |
| Magnesium chloride |  |  |  |
| Iron sulfate |  |  |  |
| Calcium sulfate |  |  |  |

**2 – The first part of the name of a salt comes from the alkali, usually from the metal in the alkali. The second part of the name of a salt comes from the acid.**

**Complete the following neutralisation reactions predicting the salt produced in each reaction. The first one has been done for you.**

Hydro***chloric*** acid + ***magnesium*** hydroxide ***Magnesium chloride***

b) Hydrochloric acid + sodium carbonate

c) Sulfuric acid + calcium carbonate

d) Sulfuric acid + sodium hydroxide

e) Nitric acid + calcium hydroxide

**Choose the correct answer.**

* Identify all the observations from those below that would suggest that a chemicalreaction had taken place.
1. Bubbles of gas
2. There are no new products
3. The reaction can be reversed
4. Colour change
5. The test tube feels warmer
* Metals like gold and silver are used to make jewelry because they don’t react with acids. They are considered precious metals because they are
1. Highly reactive b- reactive

c- unreactive d- expensive

* The two products of the reaction of zinc with hydrochloric acid are:
1. H2O + H+ b- H2O + ZnCl2

c- ZnCl2 + H+ d- H+  + HCl

* When iron reacts with sulfuric acid. The products are a salt and hydrogen.

The best equation to show this reaction is

1. sulfuric acid + iron → iron chloride + hydrogen
2. sulfuric acid + iron → iron sulfate + hydrogen
3. sulfuric acid + iron → iron sulfate + water
4. sulfuric acid + iron → iron chloride + water