**Name:** \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ **worksheet:** Acids & Alkalis

**Date:** \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ **Grade:** 6CS

**Objectives:**

* Compare the properties of acids and alkalis and give examples.
* Know the pH of acidic, alkaline and neutral solutions.

***Question One:*** *Compare and contrast acids and alkalis by completing the following table.*

|  |  |  |
| --- | --- | --- |
|  | ***Acids*** | ***Alkalis*** |
| ***pH range*** |  |  |
| ***What to look for in chemical formula*** |  |  |
| ***Production of ions*** |  |  |
| ***Taste*** |  |  |
| ***Examples*** |  |  |

***Question Two:*** *Classify the following examples as acids, bases, or salts.*

HBr \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ KCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Mg(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ H3PO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

HClO \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ KNO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Al(OH)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ HFO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Ba(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ CaCO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

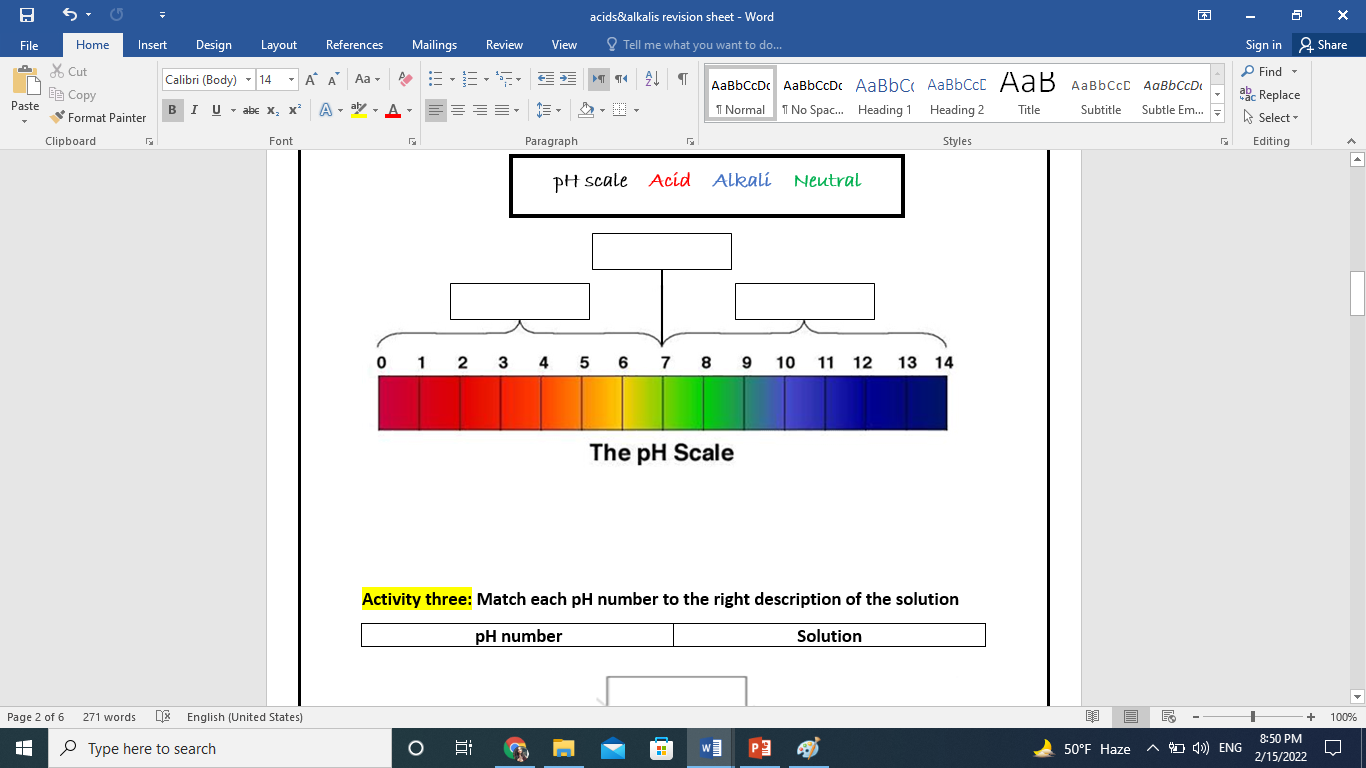
***Question Three:*** *Name the following acids and alkalis.*

HCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ HNO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

H2SO4\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ NaOH \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

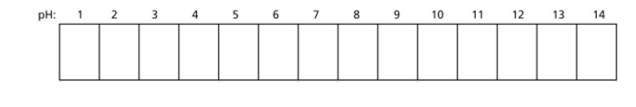
KOH \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Mg(OH)2\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

***Question Four:*** *Use the pH scale to answer the following questions.*



1. Fill in the boxes with the correct words.
2. How would you describe a substance with a pH of 6? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. Sodium hydroxide is a very strong base so it would have a pH of \_\_\_\_\_\_\_\_\_\_\_.
4. What pH would you expect a strong acid to have? \_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. Neutral substances would have a pH of \_\_\_\_\_\_\_\_\_\_\_\_\_\_.

***Question Five:***

The pH scale demonstrates how strong an acid or an alkali is. The colors on a pH color chart show the color that universal indicator turns with acids and alkalis of different strengths.

A B C D E

\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_

a) *Color the pH color chart below to show what color universal indicator turns with different strengths of acids and alkalis.*

b) *Identify the labels A to E, choosing from the words below:*

***strong acid weak acid strong alkali weak alkali neutral***

***Question Six:***

*Fill in the table below with the expected results of using* ***litmus paper.***

|  |  |  |  |
| --- | --- | --- | --- |
|  | Acidic solution | Neutral solution | Alkaline solution |
| Blue litmus paper |  |  |  |
| Red litmus paper |  |  |  |

***Question Seven:*** Complete the following sentences.

1. A neutralization reaction will always produce \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
2. A solution is neutral at pH\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
3. An alkali is able to “cancel” out an acid. The chemical name for cancelling out the acid is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

***Question Eight:*** Complete the following neutralization reactions.

\_\_\_\_\_\_\_\_\_ + KOH KBr + \_\_\_\_\_\_\_\_\_

Potassium must come from the base. Therefore, the base must be KOH.

Bromine must come from the acid. Therefore, the acid must be HBr.

HCl + \_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_ + H2O