ENDOTHERMIC & EXOTHERMIC

OBJECTIVES:

To explain the terms endothermic and exothermic

To give examples of reactions for each

KEY WORDS:

ENDOTHERMIC

EXOTHERMIC

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Chemical reactions can either release energy to their surroundings, ***EXOTHERMIC***, or energy can be transferred to them from the surroundings, ***ENDOTHERMIC***.

https://www.youtube.com/watch?v=0cUK4jcAEaU&t=190s

EXAMPLES:

Exothermic: Burning magnesium/ reactions of metals with acids

Endothermic: Photosynthesis

How do you know what type of reaction is happening?

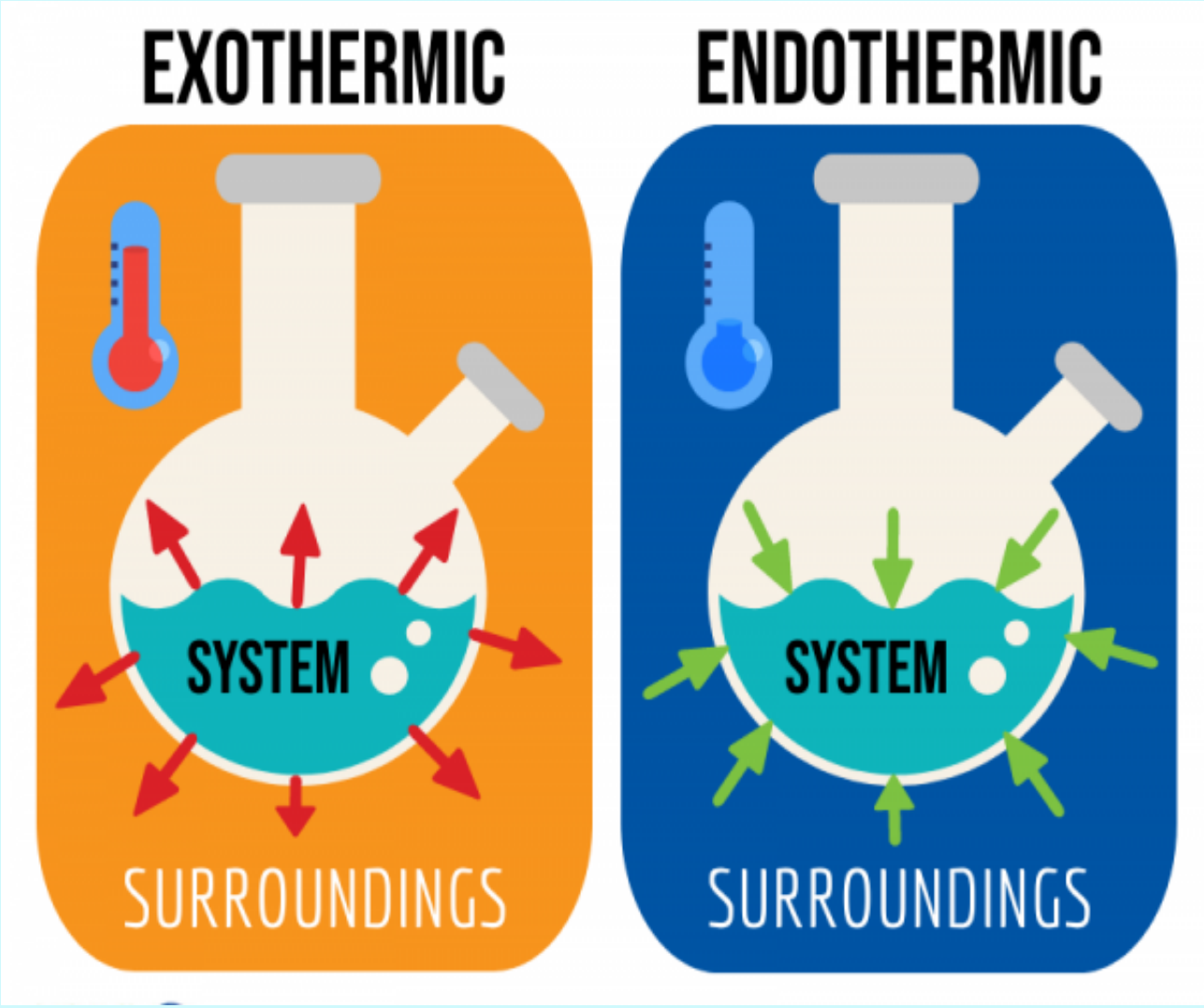
One way to do this is by looking at the system and surroundings of a reaction.
The system is where the reaction takes place, and the surroundings are the area around the system.

To determine if a reaction is exothermic or endothermic, you could *measure the temperature change of the system or its surroundings.*
you simply measure the temperature of a reaction before and after it is completed.

Differences Between Endothermic and Exothermic Reactions

Endothermic Reactions: Endothermic reactions are chemical reactions that absorb heat energy from the surrounding.
Exothermic Reactions: Exothermic reactions are chemical reactions that release heat energy to the surrounding.

Endothermic Reactions: The temperature decrease.
Exothermic Reactions: The temperature increase.


EXOTHERMIC
SYSTEM
SURROUNDINGS
ENDOTHERMIC
SYSTEM
SURROUNDINGS