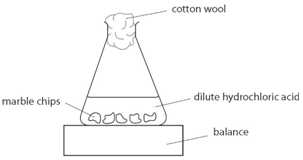
**Question one:**

Small lumps of marble chips were added to dilute hydrochloric acid in the apparatus shown in the diagram.

The balance reading was noted as soon as the lumps were added, and again every 20 seconds. The experiment was repeated by using different sizes of the marble chips.

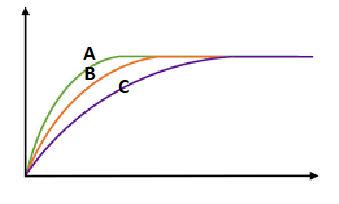
1. **Write the word equation for this reaction.**

**Caco3+hcl………cacl2+h2o+co2**

1. **Suggest why a cotton wool was placed in the mouth of the conical flask.**

**To prevent any acid spraying out**

**The results of the investigation were plotted on the graph as shown.**



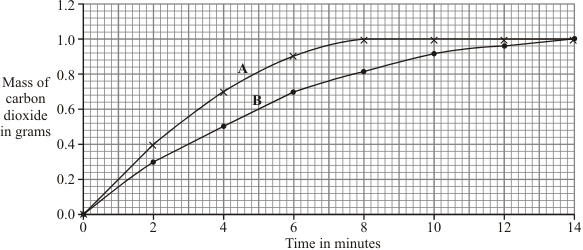
**Mass of co2**

**Time(seconds)**

1. **Label the X axis and the Y-axis**
2. **Label the marble chips with the smallest surface area. (A)**
3. **Decide if the following statements are true or false:**
4. **Line A represents the results of the large chunks reaction. ……F…..**
5. **All the three experiments ended at the same time and caused the same mass loss. …F…….**
6. **The reactions are the fastest at the beginning. …T…….**
7. **The three lines turned flat because the students stopped collecting data. ……F….**

**Question two:**

**Two 10g samples of marble of different sizes, A and B, were each reacted with 50 cm3 of diluted hydrochloric acid. The mass of carbon dioxide formed in each reaction was recorded and plotted to produce the graph below.**



1. **Identify:**

* The dependent variable …time………………………………
* The independent variable …marble size ………………………………….

1. **Fil in the table below with the missing results.**

|  |  |  |
| --- | --- | --- |
| **Time/ min** | **Sample** | **Mass of gas produced/g** |
| **4** | **B** | **0.5** |
| **8** | **B** | **0.8** |
| **14** | **B** | **1.0** |
| **2** | **A** | **0.4** |

1. **How long did it take sample A to finish the reaction?** …8min……………….
2. **Explain the results of the investigation, in terms of particles and collisions.** As the marble size increases the time taken(min) increases so, for a chemical reaction to occur, the reacting particles must collide with one another.

The rate of the reaction depends on the frequency of collisions.

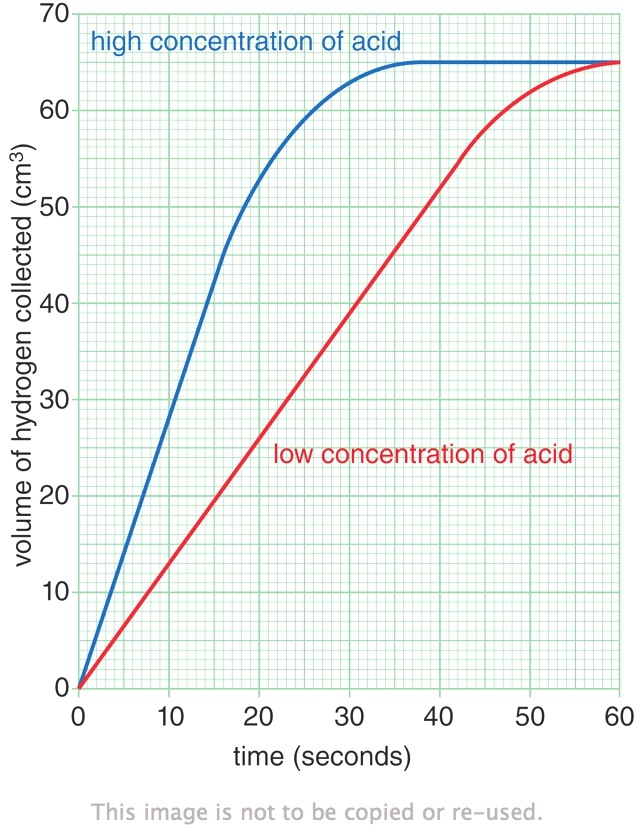
**Question three:**

**Some students were investigating how fast hydrogen gas is released in the reaction between magnesium and hydrochloric acid.**

**They used 0.1 g of magnesium ribbon with specific volume of diluted acid.**

**Next, they repeated the experiment using magnesium ribbon with the same volume of concentrated acid.**

**Their results are shown on the graph.**



1. **How do you explain that the same volume of gas was given out in both experiments?**

as shown in the graph both high and low concentrated acid collected the same amount of the volume of hydrogen collected (65cm3)

1. **The faster reaction was caused by using a concentrated acid. Explain, in terms of particles and collisions, why a higher concentration acid causes a faster rate of reaction.**

If the concentration of reactants is increased, there are more reactant particles moving together. There will be more collisions and so the reaction rate is increased. The higher the concentration of reactants, the faster the rate of a reaction will be.

1. **Why do you think the reaction with the diluted acid got slower between 45-55 sec?**

.. Reactant concentration decreases with time as the reactants are converted to products between (45-55 seconds)