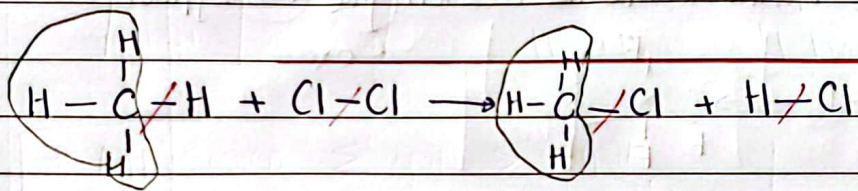
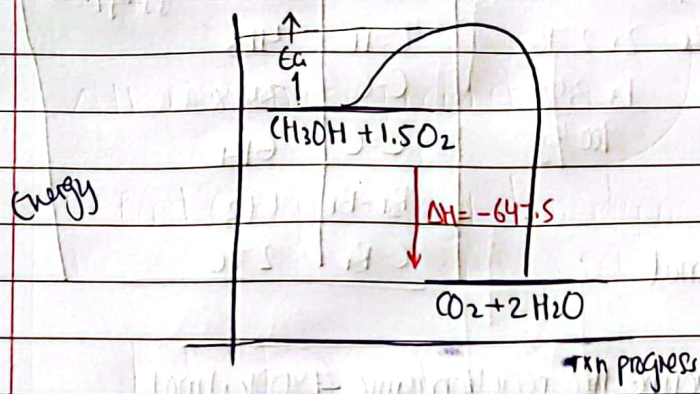


$$\Delta H = \epsilon_{\text{input}} - \epsilon_{\text{output}}$$

$$= 2802.5 - 3450$$

$$= -647.5 \text{ kJ/mol}$$

~~exo~~



Bond Broken

Bond Formed

1 x C-H	413	1 x C-Cl	328
1 x Cl-Cl	242	1 x H-Cl	431
	655 kJ		957 kJ

Bond	Bond Energy (kJ/mole)
C-H	413
Cl-Cl	242
H-Cl	431
C-Cl	328

$$\rightarrow \Delta H = 655 - 957 = -302 \text{ kJ/mol}$$

~~exo~~

