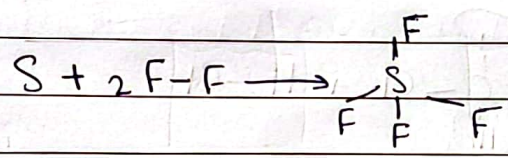


Bond Broken		Bond Formed		Bond	Bond Energy kJ/mole
Br-Br	193	2 x C-Br	2 x 276	C-H	413
C=C	614	C-C	1 x 348	C-C	348
	807 kJ		900 kJ	C=C	614
				Br-Br	193
				C-Br	276

$$\Delta H = 807 - 900 = -93 \text{ kJ/mol}$$

• when sulfur react with fluorine the reaction give ~~-780~~  $-780 \text{ kJ/mol}$ .



→ if the bond energy of F-F is 160 kJ/mol. find the bond energy of S-F? And draw an energy level of diagram:

