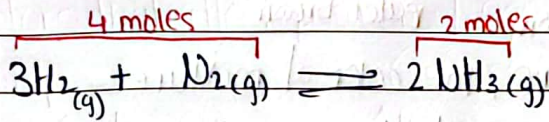


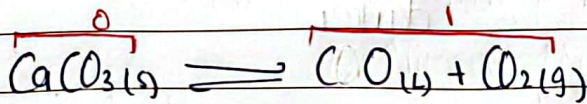
## ② Pressure

→ As the pressure increases, the equil shift to the side with low pressure. with less gas moles  
→ " " decreases, " " " " with high pressure with high gas moles.



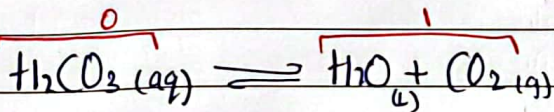
↑ Pressure, shift to the side with less gas moles ↑ NH<sub>3</sub>

↓ Pressure, " " " " " " more " " " "

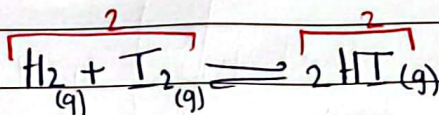


↓ pressure, shift to the side with more gas mole

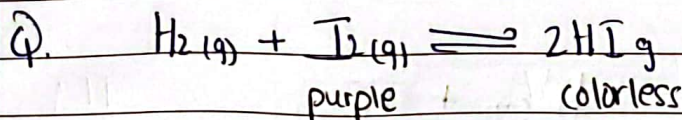
↑ " " " " " " less " " " "



↓ pressure, shift to the side with more gas moles.



⇒ Changing the pressure has no effect on the position of equilibrium since both sides has the same no. of gas moles.



the equil doesn't affect by increasing the pressure

→ why by increasing the pressure the mixture becomes more purple.

