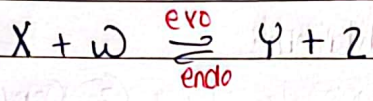


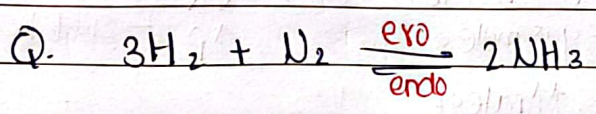
$$\Delta H = +ve$$

↑ Temp ↑↑ rate of forward
 ↑ " " backward
 ↓ A ↓ B ↑ C ↑ D



$$\Delta H = -ve$$

↑ Temp ↑ rate of forward) shift to endo
 ↑↑ " " backward)
 ↑ X ↑ W ↓ Y ↓ Z

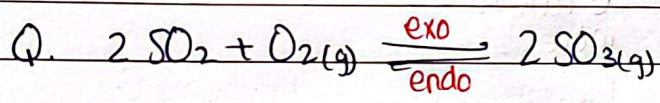


$$\Delta H = -ve$$

→ to produce more yield of NH_3 -

→ we must use low temp.

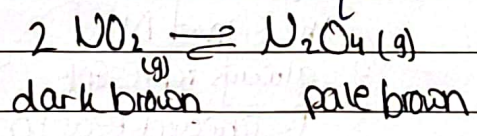
→ to favour the forward reaction which is exothermic.



$$\Delta H = -ve$$

	rate of forward	rate of back	% SO_3
↑ Temp	increase	increase	decrease
↓ Temp	decrease	decrease	increase.

• Mixture of NO_2 and N_2O_4 at equil in a sealed mixture.



• if we put this sealed tube in a cold water bath, the mixture becomes paler? why?

→ because forward reaction is exothermic enhanced by cooling.