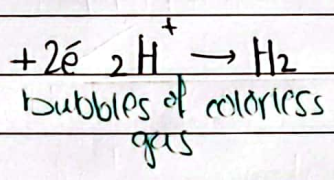
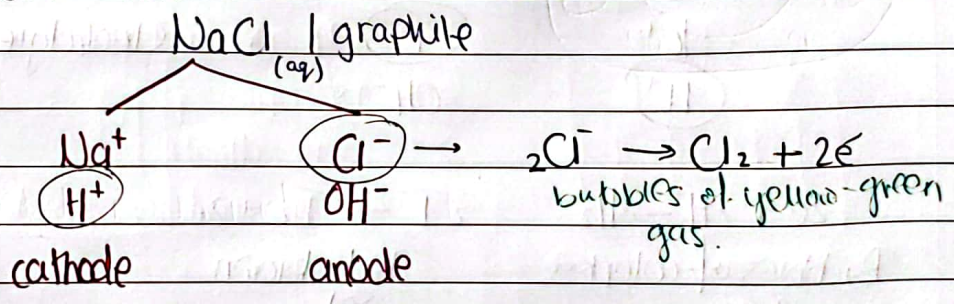
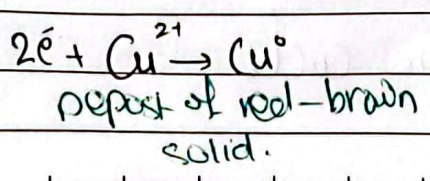
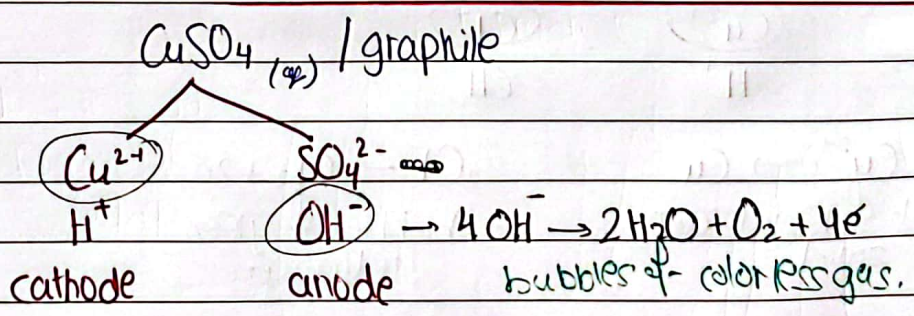


	anode	
X	$K^+$ $Na^+$ $Li^+$ $Ca^{2+}$ $Mg^{2+}$ $Al^{3+}$ $Zn^{2+}$ $Fe^{2+,3+}$ $Pb^{2+}$	Always $OH^-$ except concentrated halide ( $Cl^-$ , $Br^-$ , $I^-$ )  * when the halide oxidise $2Cl^- \rightarrow Cl_2 + 2e^-$  * when $OH^-$ oxidise $4OH^- \rightarrow 2H_2O + O_2 + 4e^-$
	$2e^- + H^+ \rightarrow H_2$ $Cu^{+}, +2$ $Ag^+$ $Au^+$	

Ex: - concentrated



Electrolyte:  $NaOH$  (ferr)



Electrolyte:  $H_2SO_4$