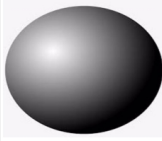


History Of Atomic Model

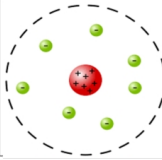
Atomic model : John Dalton



1. All matter is made up of tiny, indivisible particles called atoms.
2. All atoms of a specific element are identical in mass, size, and other properties. However, atoms of different elements exhibit different properties and vary in mass and size.

1803 CE

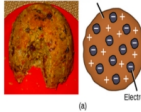
Nuclear model: Ernest Rutherford



1. Atoms are mostly empty spaces.
2. Most of the mass is concentrated in the center of the atom. This tiny, dense, positively charged core is called a nucleus.
3. Electrons are located outside the nucleus.

1911 CE

Plum-pudding model: J.J. Thomson



Positively charged matter

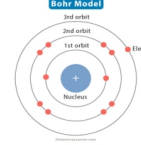
Electron

(a)

The atom is made out of a sphere of positive charges with negatively charged electrons embedded in it.

1904 CE

Planetary model: Niels Bohr



Bohr Model

3rd orbit

2nd orbit

1st orbit

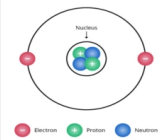
Electron

Nucleus

Electrons orbit the nucleus in orbits that have specific size and energy.

1913 CE

James Chadwick Atomic Theory



Electron

Proton

Neutron

Confirmed the existence of neutrons, Nuclei contain both neutrons and protons

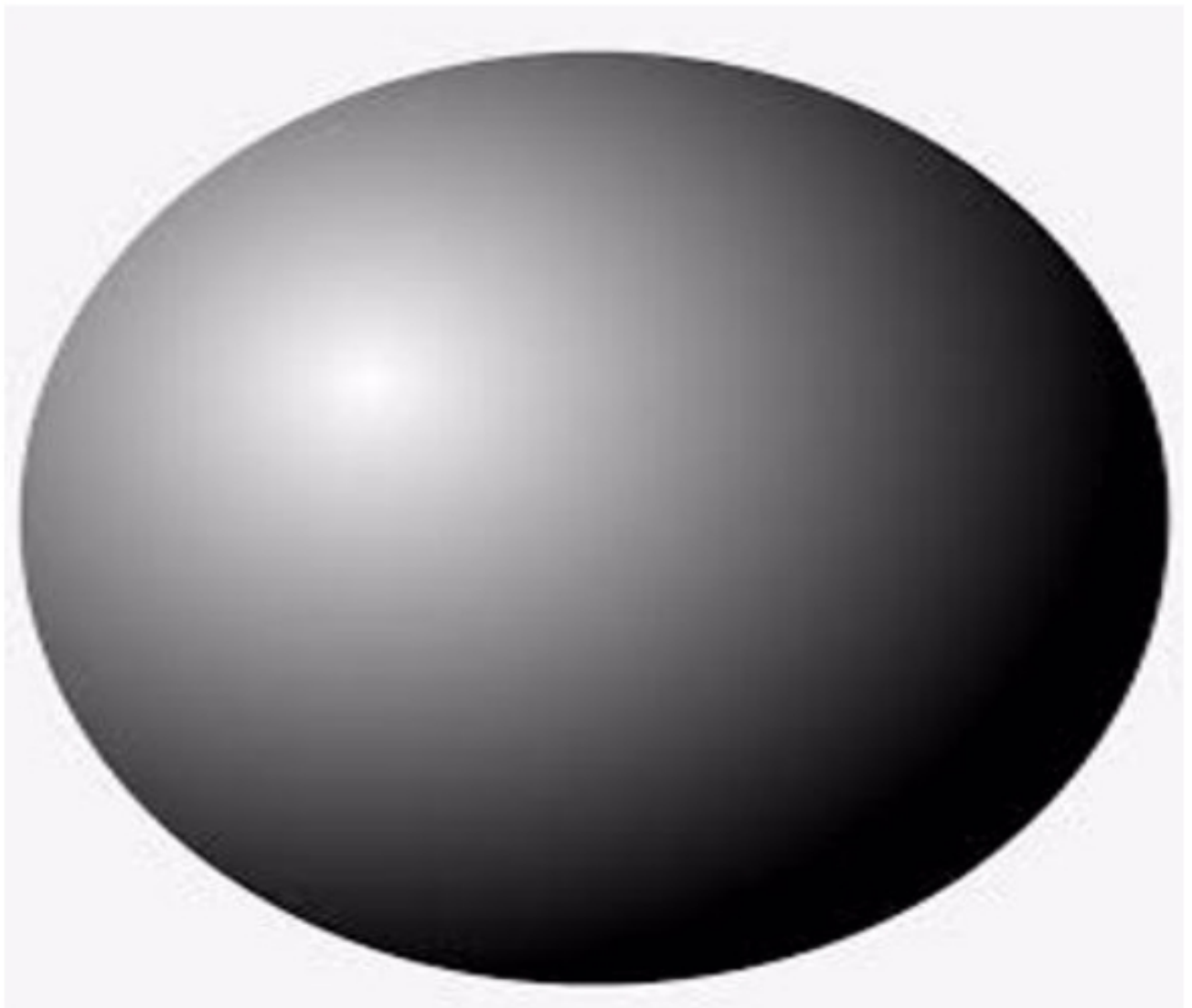
1932 CE

Legend

10⁷ Years and 364 Days

// Time Break

Atomic model : John Dalton



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Plum-pudding model: J.J. Thomson



Positively charged matter

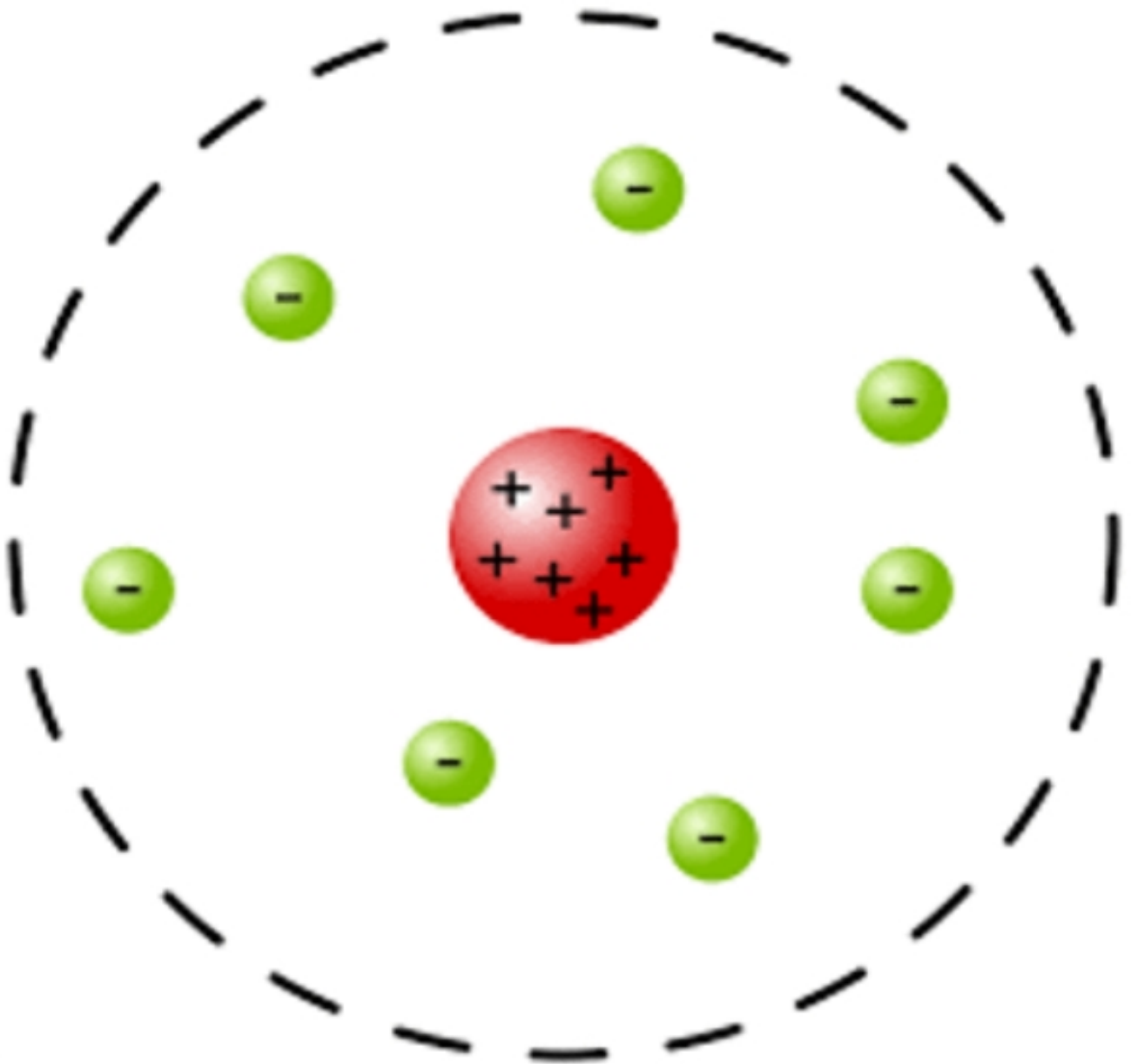


Electron

(a)

The atom is made out of a sphere of positive charges with negatively charged electron embedded in it.

Nuclear model: Ernest Rutherford



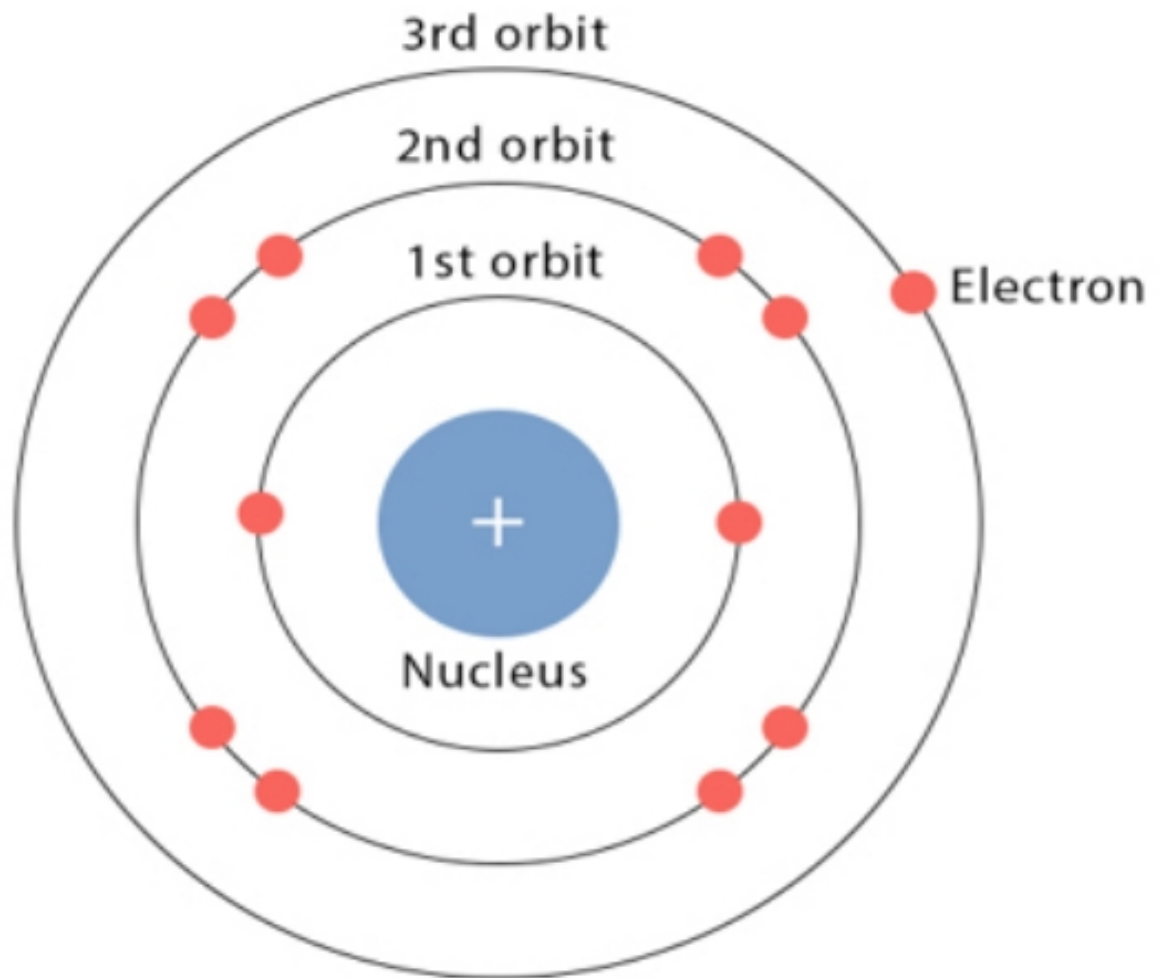
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Planetary model: Niels Bohr

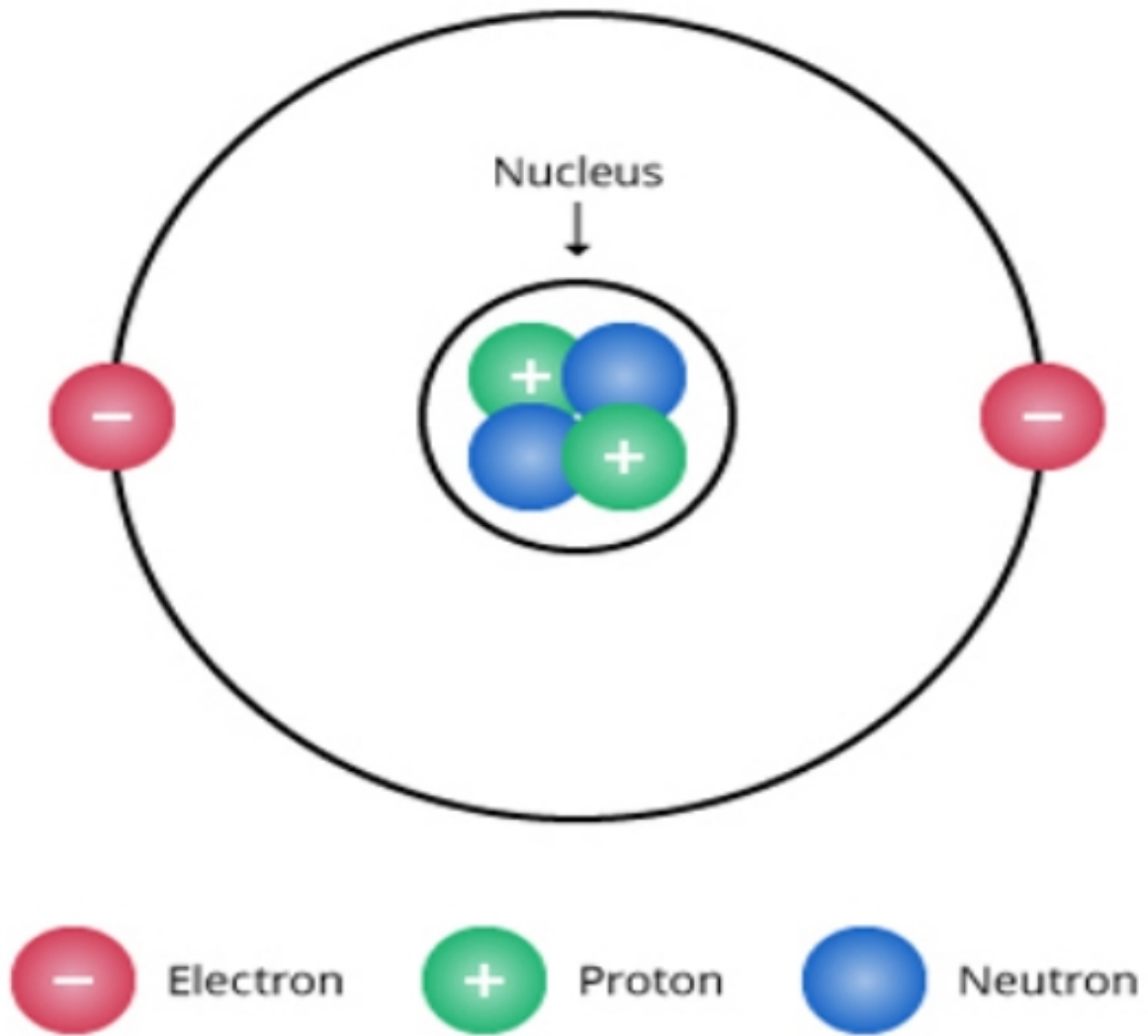
Bohr Model



ChemistryLearner.com

Electrons orbit the nucleus in orbits that have specific size and energy.

James Chadwick Atomic Theory



Confirmed the existence of neutrons, Nuclei contain both neutrons and protons